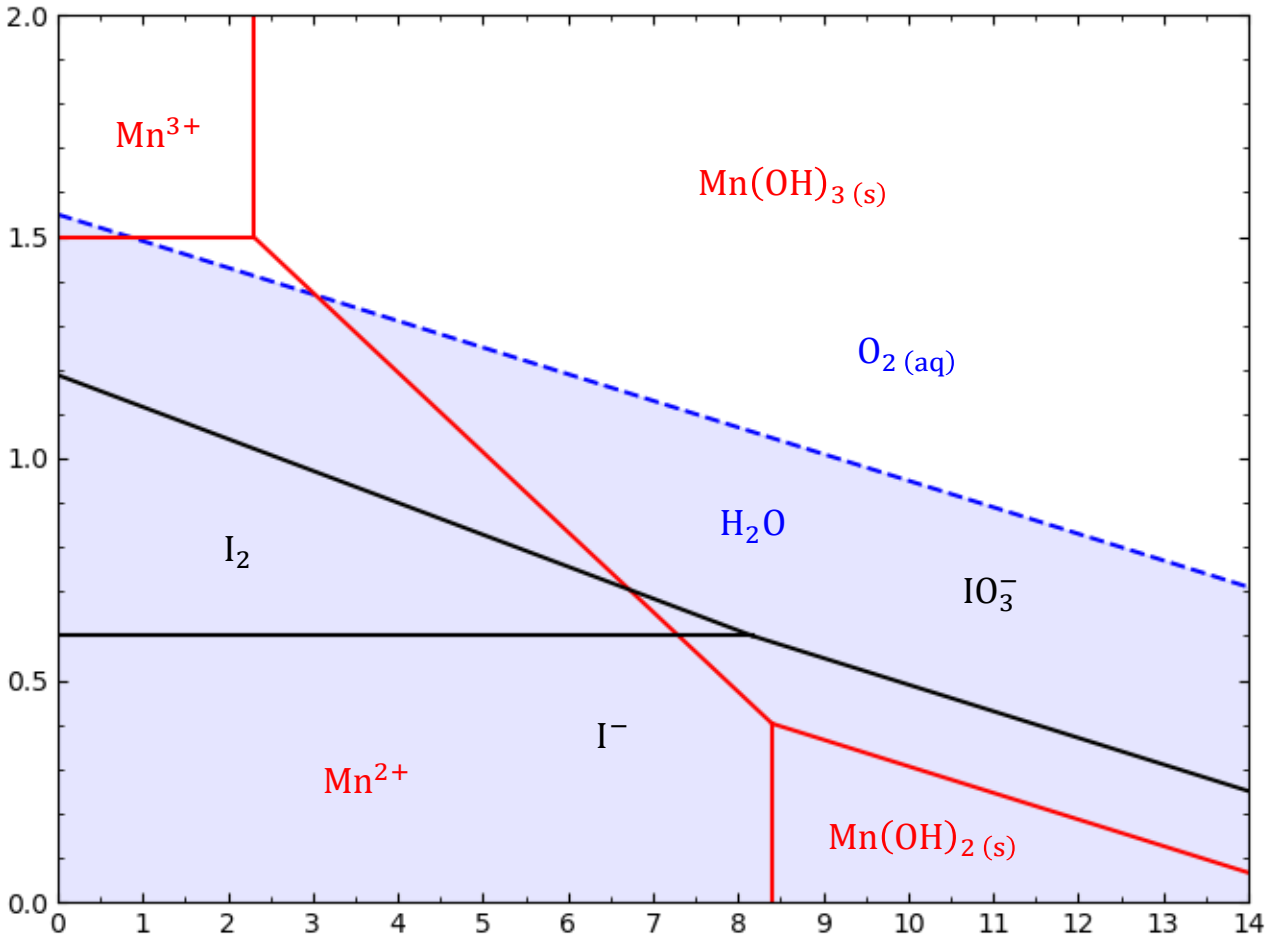
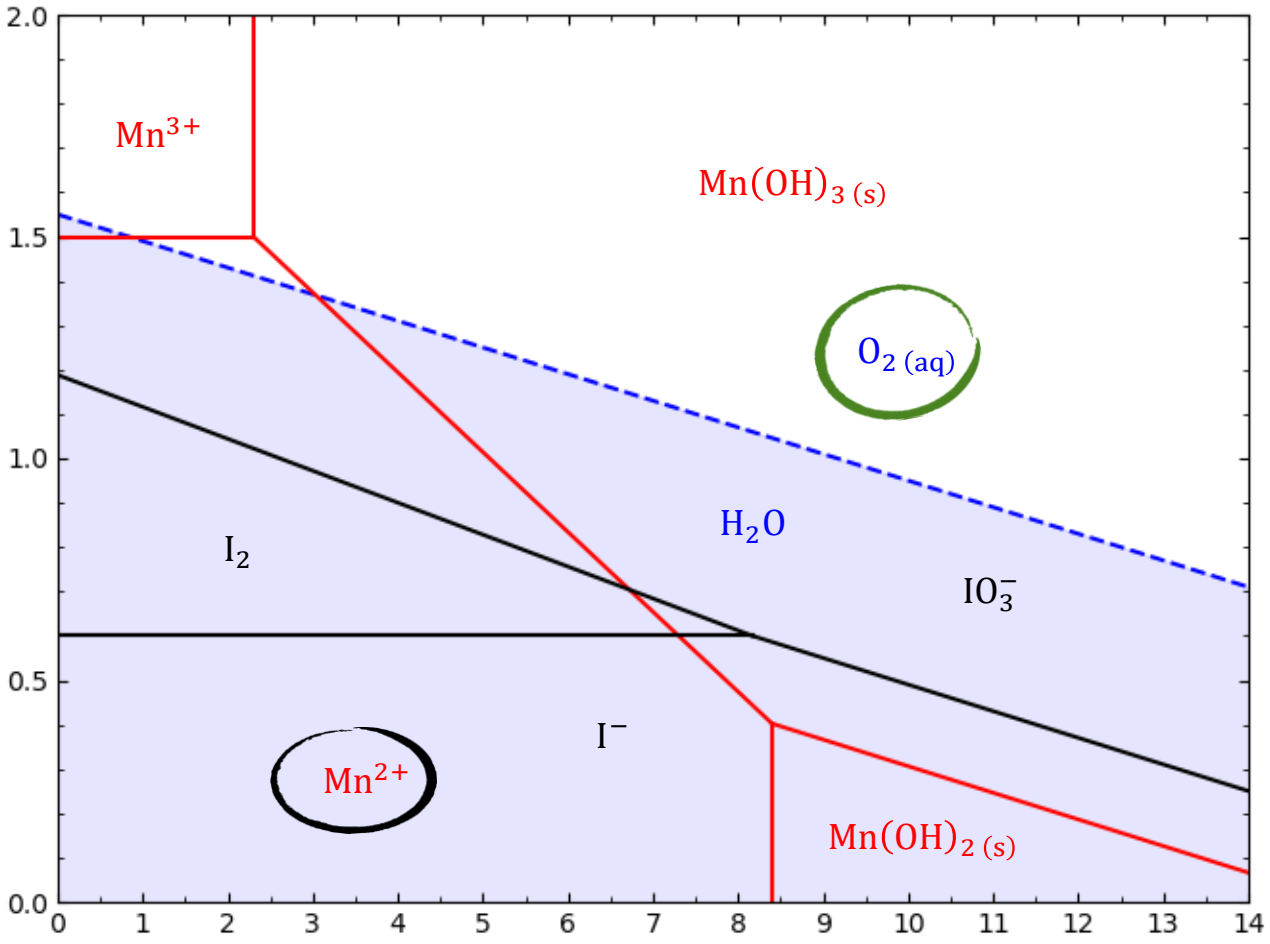
Diagramme E-pH de l'eau, de l'iode et du manganèse

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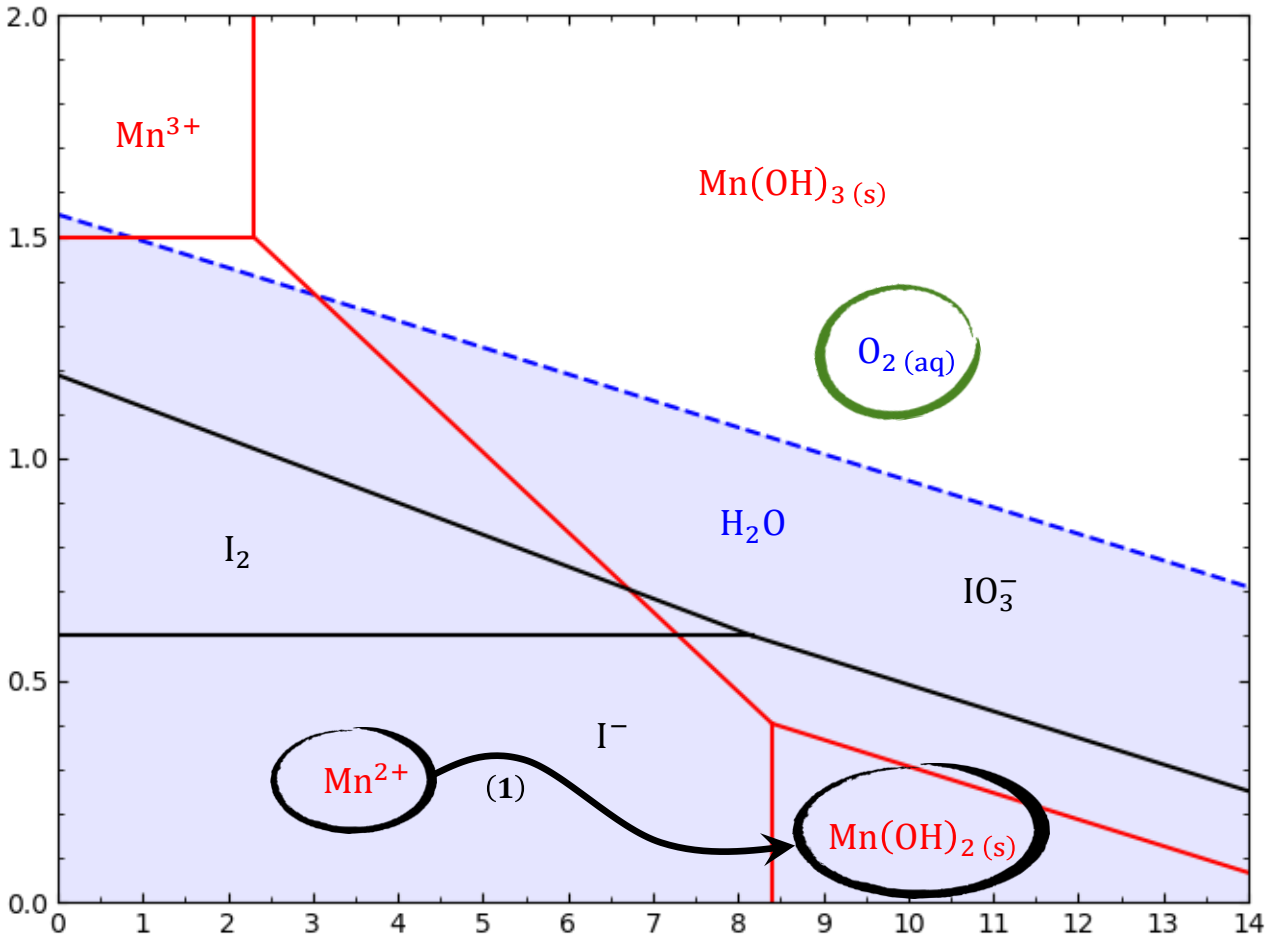
Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$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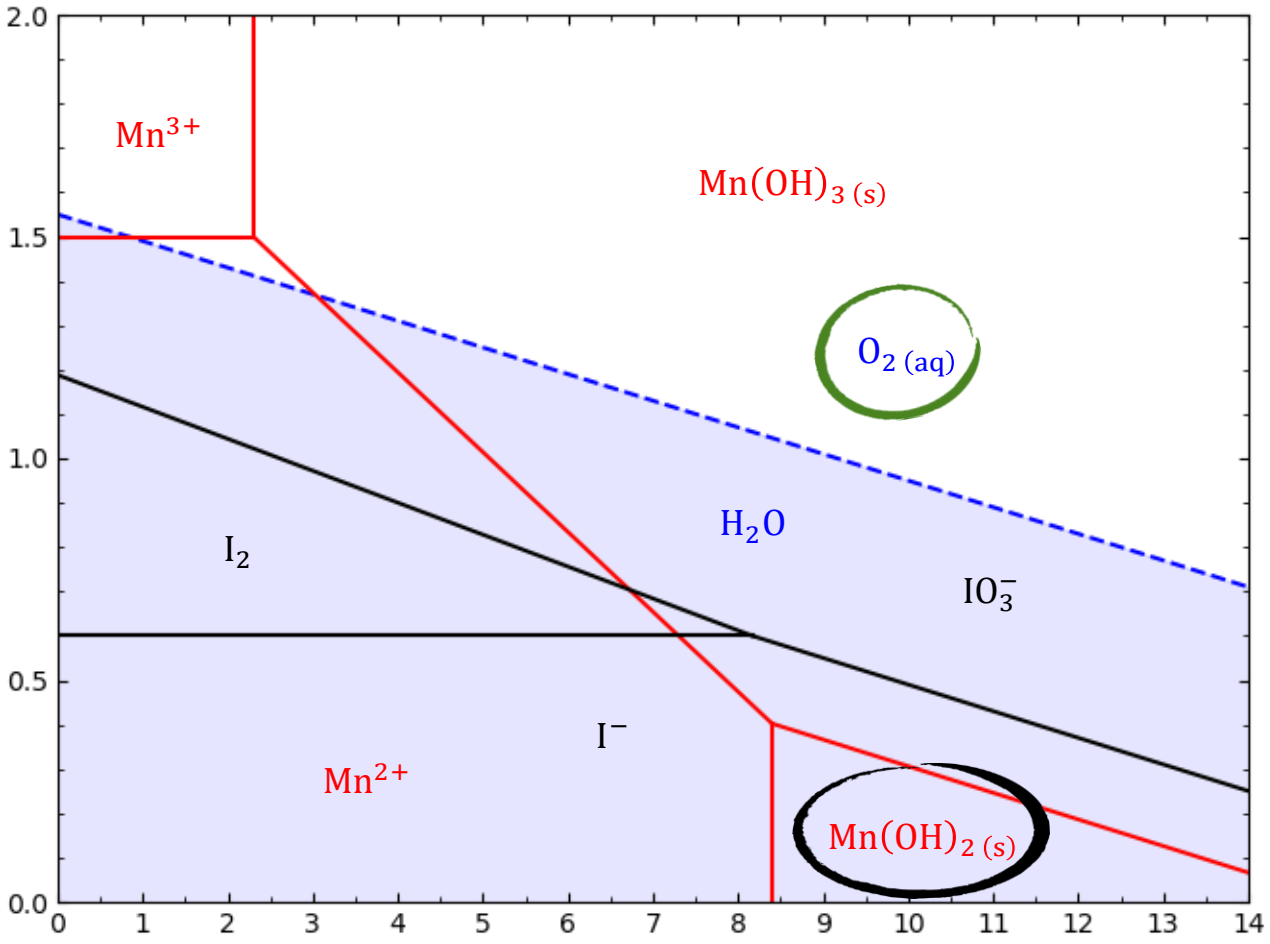
Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
(1)
$Mn(OH)_{2\,(s)}$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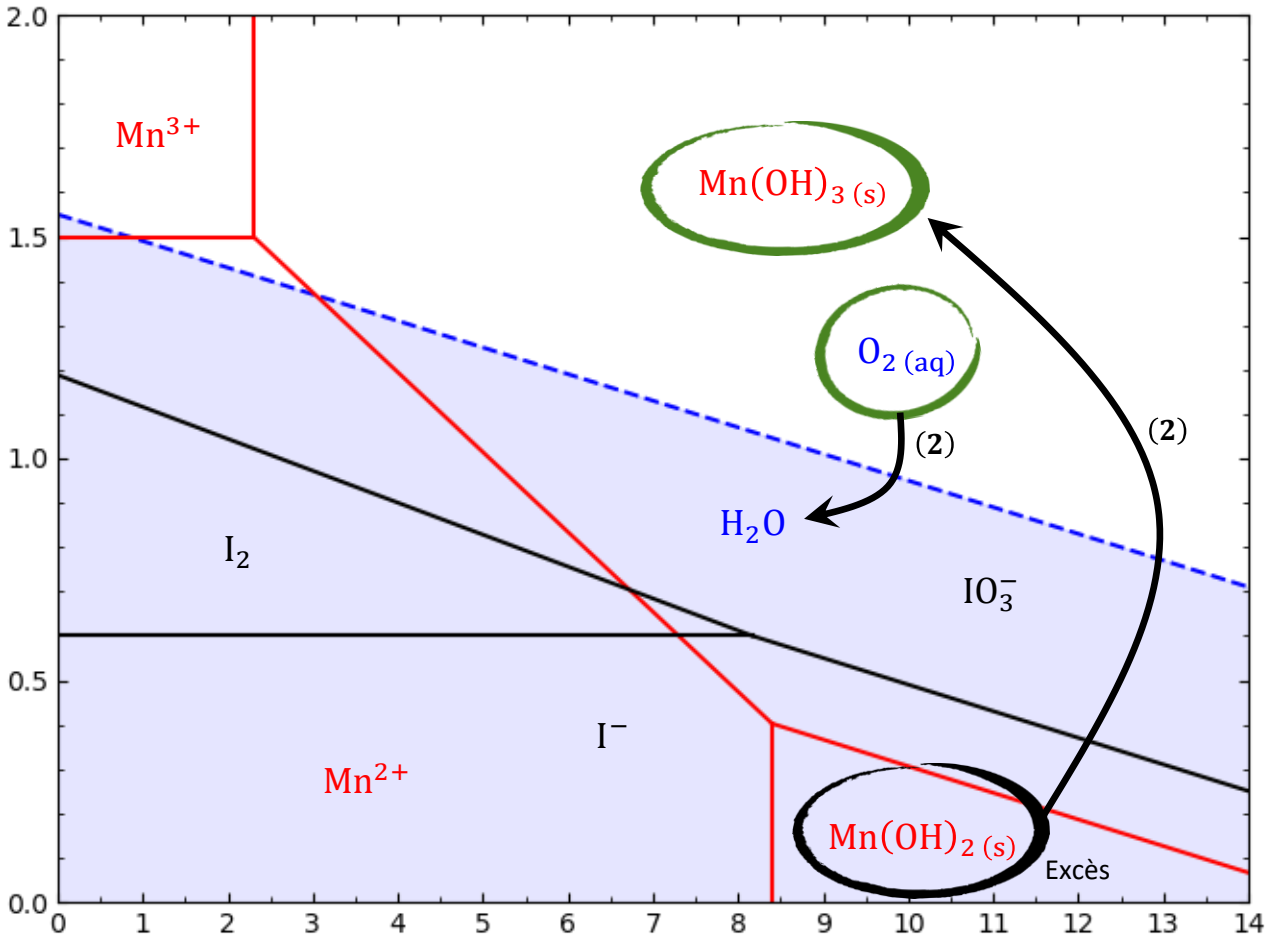
Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$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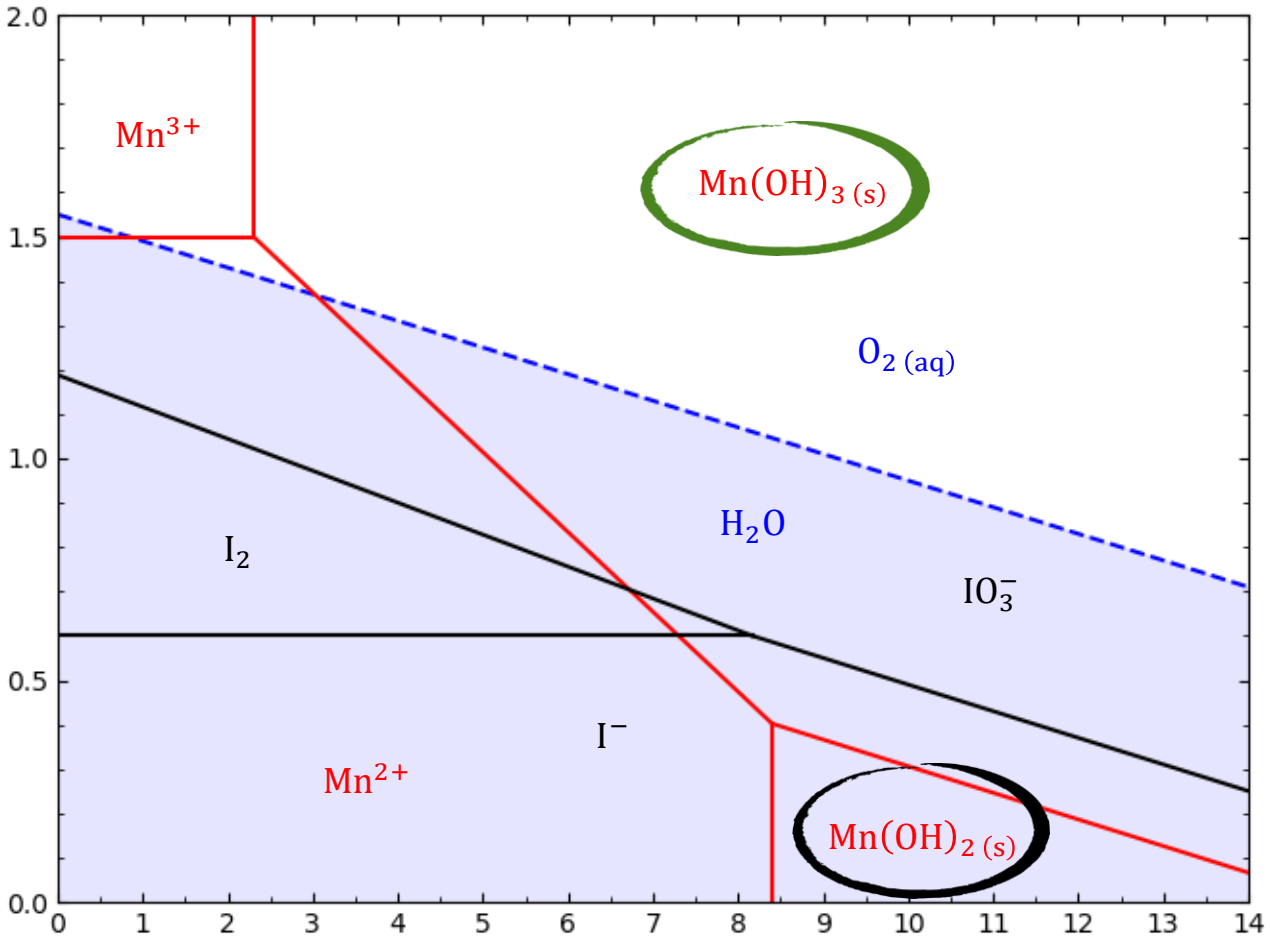
Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
(2)
(2)
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$
Excès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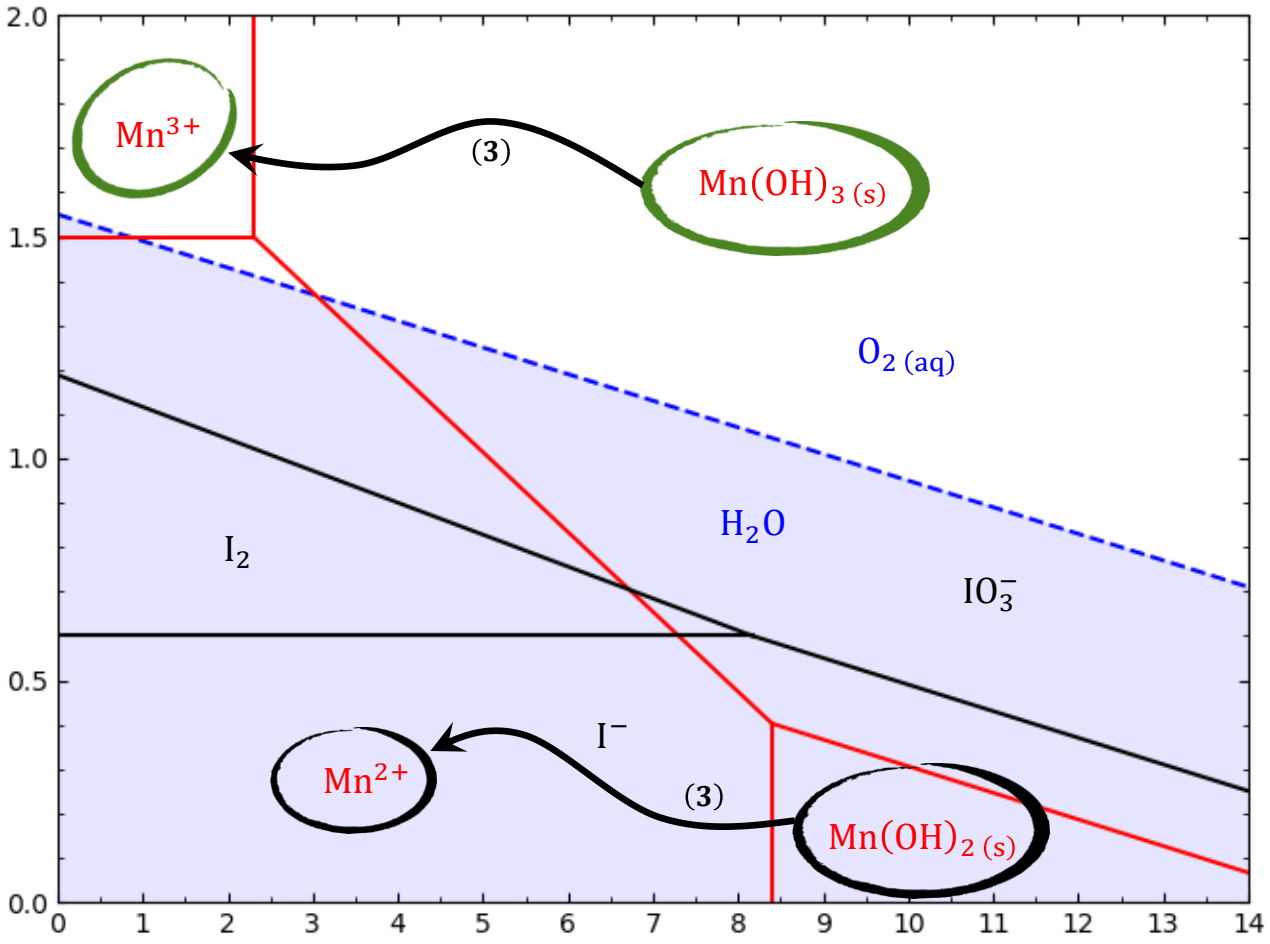
Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$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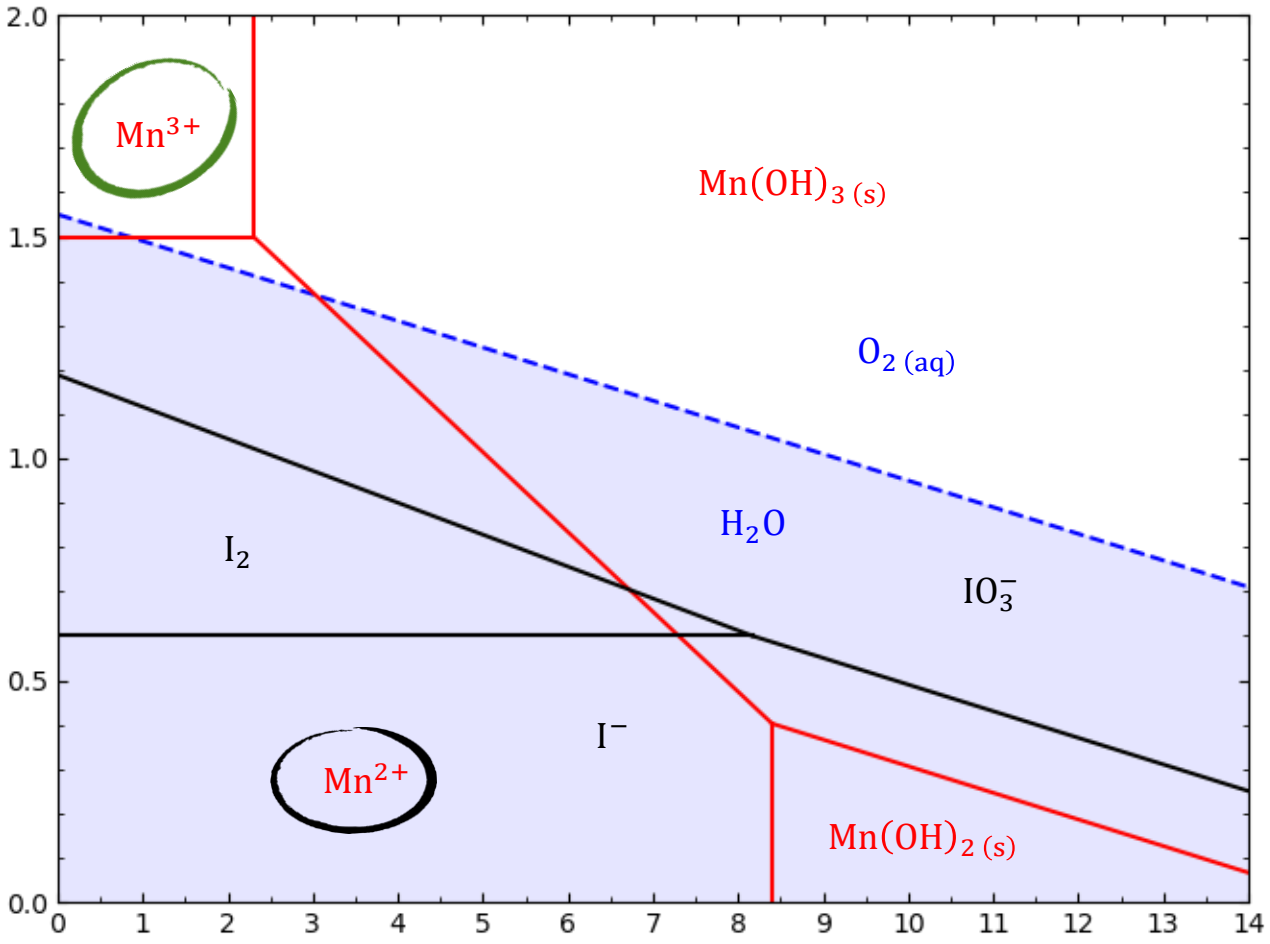
Mn^{3+}
(3)
$Mn(OH)_{3\ (s)}$
$O_{2\ (aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
(3)
$Mn(OH)_{2\ (s)}$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Mn^{3+}
$Mn(OH)_{3\,(s)}$
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
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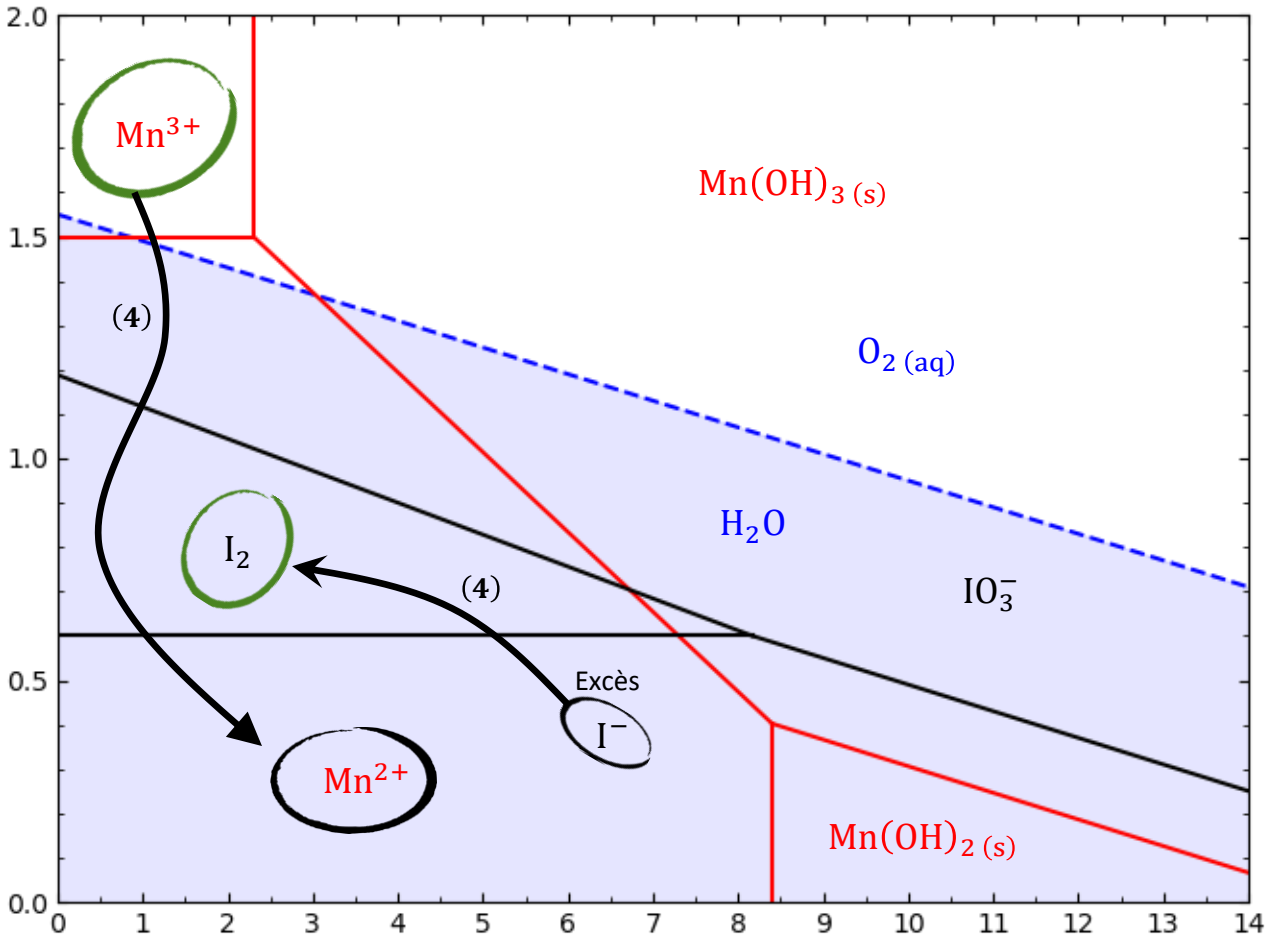
Mn^{2+}
$Mn(OH)_{2\,(s)}$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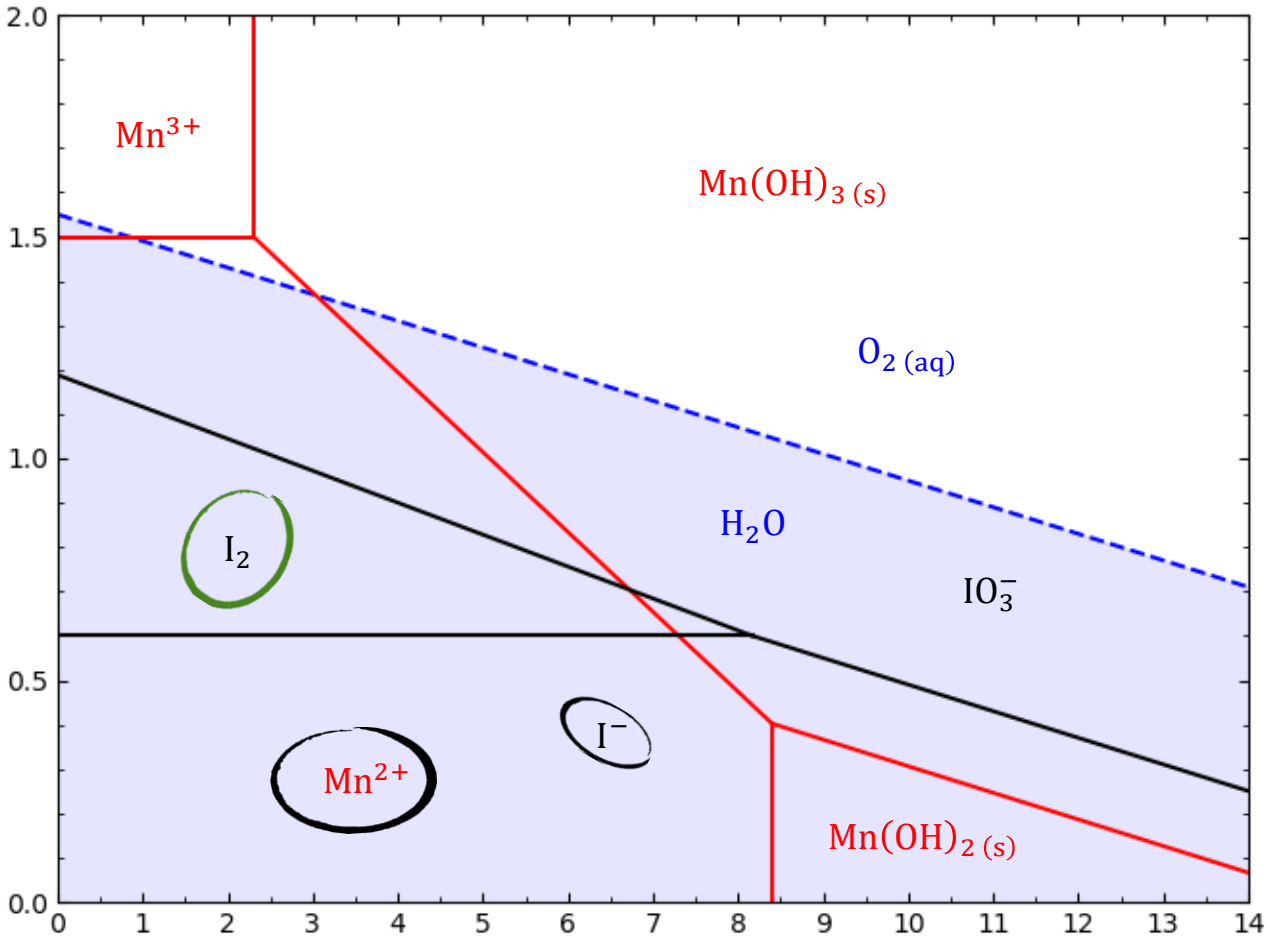
Mn^{3+}
$Mn(OH)_{3\,(s)}$
(4)
$O_{2\,(aq)}$
H_2O
I_2
(4)
IO_3^-
Excès
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$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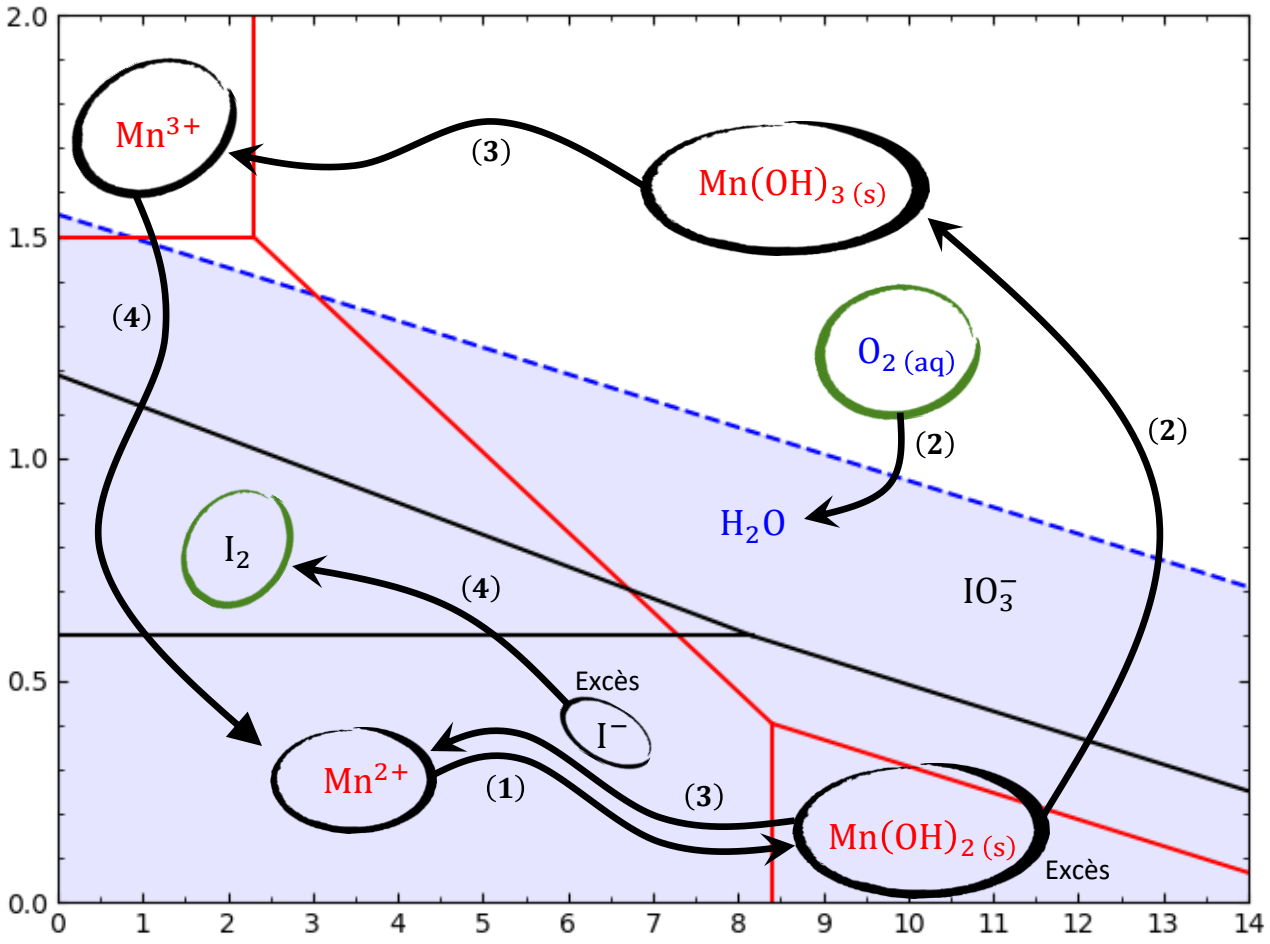
$O_{2\,(aq)}$
H_2O
I_2
IO_3^-
I^-
Mn^{2+}
$Mn(OH)_{2\,(s)}$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Mn^{3+}
(3)
$Mn(OH)_{3\,(s)}$
(4)
$O_{2\,(aq)}$
(2)
(2)
H_2O
I_2
(4)
IO_3^-
Excès
I^-
Mn^{2+}
(1)
(3)
$Mn(OH)_{2\,(s)}$
Excès